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reactants-Mass to Mole Conversions-  
Multiply the grams by the g/mol ratio  
in the reference table-Ex: 15g of C \*  
 $1/12.011 = 1.3$  grams-Molar Ratio-  
Find the ratio of between 2  
substances in a formula-Ex:  $\text{CuCl} + \text{Zn}$   
 $\text{ZnCl} + \text{Cu}$ -The molar mass formula  
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mass in grams of 1 mole of any  
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Assessment 36. a. Two formula units  $\text{KClO}_3$  decompose to form two formula units  $\text{KCl}$  and three molecules  $\text{O}_2$ . b.

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Stoichiometry problem, the given  
quantity is first converted to \_\_\_\_\_.

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, 154.21 g/mol (  $6 \cdot 12 + 5 \cdot 1 = 77$  :

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$154/77 = 2$ ) C<sub>12</sub> H<sub>10</sub> C<sub>3</sub> H<sub>6</sub> O<sub>2</sub>,  
222.24 g/mol ( $12 \cdot 3 + 6 \cdot 1 + 2 \cdot 16 =$   
77:  $222/74 = 3$ ) C<sub>9</sub> H<sub>18</sub> O<sub>6</sub> 3.

Empirical Formula from Percentage  
Composition (don ' t round until the  
very end!!!!) A compound is composed  
of 12.590% hydrogen, 37.404%  
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backpacker.com.br In the equation  
that follows each problem, write in the  
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